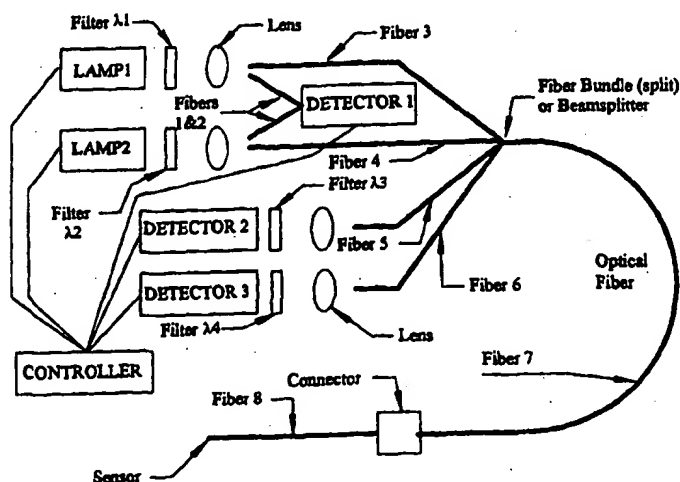




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(21) International Application Number: <b>PCT/US00/00800</b> (22) International Filing Date: <b>11 January 2000 (11.01.00)</b> (30) Priority Data: 09/228,819          11 January 1999 (11.01.99)          US (71) Applicant: <b>LIGHTSENSE CORPORATION [US/US]; Suite B, 23452 Peralta Drive, Laguna Hills, CA 92653 (US).</b> (72) Inventors: <b>RING, Lawrence, S.; 898 Hillcrest Drive, Laguna Beach, CA 92651 (US). ZOVAL, Jim, V.; 21361 Kirkwall Lane, Lake Forest, CA 92630 (US).</b> (74) Agents: <b>KLEIN, Howard, J. et al.; Klein &amp; Szekeres, LLP, Suite 700, 4199 Campus Drive, Irvine, CA 92612 (US).</b>	(81) Designated States: <b>AE, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, CA, CH, CN, CR, CU, CZ, DE, DK, DM, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, TZ, UA, UG, UZ, VN, YU, ZA, ZW, ARIPO patent (GH, GM, KE, LS, MW, SD, SL, SZ, TZ, UG, ZW), Eurasian patent (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European patent (AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE), OAPI patent (BF, BJ, CF, CG, CI, CM, GA, GN, GW, ML, MR, NE, SN, TD, TG).</b>  <b>Published</b> <i>With international search report.</i> <i>Before the expiration of the time limit for amending the claims and to be republished in the event of the receipt of amendments.</i>	

(54) Title: **METHOD AND MATERIAL FOR RATIOMETRIC FLUORESCENT DETERMINATION OF ANALYTE CONCENTRATION**



## (57) Abstract

Optical sensors are used to measure the concentration of oxygen or other analytes in a sample. The ratiometric method utilizes an indicator dye which is pyranthrene, flavanthrene or a suitable derivative dispersed or immobilized in an analyte permeable matrix. The matrix is excited with light at one or more wavelengths, or fluorescence emissions of the indicator are measured at one or more wavelengths, or both excitation and emission are performed and measured at multiple wavelengths. These indicator dyes exhibit different sensitivity to quenching by the analyte at different excitation wavelengths and also at different emission wavelengths. Each emission resulting from a different excitation is divided by the measured intensity of its excitation to compensate for illumination variations. The ratio of the resulting signals for each excitation is then used for determining the analyte concentration. Alternatively, the ratio of emissions at a single excitation wavelength are used in a similar manner.

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1  
2 METHOD AND MATERIAL FOR RATIOMETRIC FLUORESCENT  
3 DETERMINATION OF ANALYTE CONCENTRATION  
4

5 BACKGROUND OF THE INVENTION

6 1. Field of the Invention:

7 The present invention relates to optical chemical sensors, and more  
8 particularly to an improved optical sensor system for use in the measurement  
9 of analytes such as oxygen, sulfur dioxide, nitric oxide, the anesthetic agent  
10 HALOTHANE, halogens, and other materials known to quench the  
11 fluorescence of polycyclic aromatic hydrocarbons (PAHs) and polycyclic  
12 heterocyclic aromatic compounds (HACs). These measurements are used to  
13 determine the concentration of analytes in gaseous and fluid mixtures. These  
14 measurements are also used for determining pressure, glucose concentration,  
15 and various other parameters by extension.

16 1. Brief Description of the Prior Art:

17 The use of fiber optic sensor systems for the measurement of parameters such  
18 as partial pressure of oxygen (PO<sub>2</sub>) and of concentration or pressure of other  
19 analytes through the use of a luminescent material is well known. These  
20 sensors have been used in medical applications (both invasive and  
21 non-invasive), chemical, biochemical, environmental, military, and industrial  
22 analysis. Immunity to electrical interference, extended calibration intervals,  
23 and small size have made them particularly important to the medical  
24 community.

25 The principle utilized in these devices for measuring concentration of  
26 analytes, such as oxygen, is that fluorescence emissions from certain dyes is  
27 quenched by the analyte in accordance with the Stern-Volmer relationship.  
28 Thus a fluorescent dye whose fluorescence is quenched by the analyte is  
29 incorporated in the sensing device as indicator. In medical applications

1 systems incorporating these sensors are often part of multipurpose catheters  
2 which measure blood pressure, provide for fluid infusion and sample  
3 withdrawal, and may include additional sensors or therapeutic tools.

4 Optical oxygen sensors may also be used for extracorporeal monitoring  
5 during heart/lung bypass procedures, dialysis, or other blood treatment  
6 procedures, for arterial line monitoring, both *ex-vivo* and *in-vivo*, tissue  
7 monitoring of peripheral perfusion, augmentation of cardiac catheters, and  
8 many others. Non-medical uses include analysis in food and beverage  
9 manufacturing, water quality monitoring, chemical and biochemical process  
10 control, environmental monitoring, and the like.

11 Optical chemical sensors are typically fabricated by mounting the  
12 indicator dye to the end of an optical fiber, window, or rod. Light is  
13 transmitted down the optical element from an excitation source, such as a  
14 laser, lamp, or an LED (light emitting diode). At the tip of the fiber, window  
15 or rod the analyte being sensed interacts with the indicator dye to change its  
16 optical properties. It is this change which is measured by measuring the light  
17 emitted by the fluorescent indicator dye. Various methods are used to separate  
18 the fluorescent emission light returning from the indicator dye from excitation  
19 light scattered within the system. For example optical filters, beam splitters,  
20 bifurcated fiber bundles, or like devices are often used to separate excitation  
21 and emission light and to measure the emission light. The results of the  
22 measurement may be displayed in any number of ways such as a digital  
23 display, and the like. The thin indicator dye matrix layer used in most optical  
24 sensors results in a relatively fast response time.

25 In an alternative configuration, the indicator dye is excited by the  
26 application of an electric field instead of light. This approach is generally  
27 referred to as electroluminescence.

28 Thus, the optical oxygen sensor is based on the principle of quenching.

1 There are numerous fluorescent compounds that exhibit a predictable change  
2 or quenching of fluorescence when exposed to oxygen or other analytes and  
3 therefore can be used as indicator dyes in the above-summarized optical  
4 chemical sensors.

5 However, due to the nature of the dyes used in the prior art and to other  
6 factors, several problems exist with prior art fiber optic chemical sensors.  
7 Some sensors are not stable. Performance varies significantly over time.  
8 Other sensors require expensive light sources or sophisticated computers.

9 The prior art has strived to eliminate some of the foregoing problems  
10 and, as a result, provided optical chemical sensors having improved  
11 performance and stability. These generally incorporate means for referencing  
12 the analyte signal to a stable independent reference signal. Optimum  
13 performance is obtained when the analyte signal and a reference signal occur  
14 within the spectrum of a single indicator dye.

15 Methods used to obtain this independent reference signal have included  
16 providing a second sensor not exposed to the analyte, the use of a separate  
17 indicator compound, the use of different chemical forms of an indicator, the  
18 use of fluorescence lifetime measurements, and the reported use of emission  
19 ratios. Even these improved prior art systems have significant drawbacks.  
20 Some of the drawbacks are low fluorescence intensity, variations in excitation  
21 light intensity, transmission changes in the optical path, concentration  
22 variations of the indicator dye(s), leaching of the dye out of the sensor,  
23 changes in the local environment at the sensor matrix such as physical damage  
24 or conformational variations, limited dye stability, power intensive light  
25 source needs, photodegradation of the dye(s) differential photodegradation of  
26 multiple dyes, complex apparatus requirements, and inadequate reference  
27 compensation. Lifetime measurement drawbacks also include the need for  
28 very high speed data collection and analysis.

1           It is also desirable to have dyes which can be excited at wavelengths  
2 above 400 nanometer (>400nm excitation) because these dyes can be used  
3 with inexpensive LED's instead of power intensive and costly lamps, lasers, or  
4 flash lamps.

5           From among the numerous prior art patent and other references which  
6 describe the above-summarized and related features of optical chemical  
7 sensors the following are noted.

8           U.S. Patent No. 5,094,959 describes an oxygen sensor using ratiometric  
9 emission measurements with single excitation at near ultraviolet wavelength.

10          U.S. Patent No. 4,476,870 describes a basic fiber optic oxygen probe.

11          U.S. Patents Nos. 5,151,603, EP,0,442,060,A2, and 5,462,880 describe  
12 ratiometric oxygen measurement systems using 2 dyes mixed together.

13          U.S. Patent No. 4,849,172 discloses a method for increasing the  
14 solubility of a polycyclic aromatic hydrocarbon in a polysiloxane polymer.

15          U.S. Patents Nos. 4,712,865, 5,015,715, and 4,746,751 disclose  
16 methods for immobilizing a polycyclic aromatic hydrocarbon on a  
17 polysiloxane polymer.

18          U.S. Patent No. 5,681,532 describes a fluorocarbon based oxygen  
19 permeable matrix for use in optical oxygen sensors along with a 2 dye system.  
20 An article in Anal. Chem. 1985, 57, pp 2556-2561, (author O. S. Wolfbeis)  
21 and U.S. Patent No. 5,155,046 describe the effect of HALOTHANE on PAH  
22 indicators.

23          The monograph "Photophysics of Aromatic Molecules", ( *John B.*  
24 *Birks*, 1969 page 444 & 504 describes that the oxygen sensor technology can  
25 also be used to measure nitric oxide, carbon tetrabromide, carbon  
26 tetrachloride, ethyl iodide and many other halogenated compounds.

27          As described in SPIE Vol. 990 1988, 116-120 the oxygen sensor  
28 technology can also be used to measure sulfur dioxide, pressure as described

1 in U.S. Patent 5,359,887, and glucose as described in U.S. Patent 5,034,189.

2 Most of the patents listed above and U.S. Patents No. 4,917,491 utilize  
3 the fluorescent dye in a fiber optic sensor arrangement. U.S. Patent No.  
4 5,629,533 utilizes the fluorescent dye in an electroluminescent analyte sensor  
5 arrangement.

6 Another pertinent background to the present invention relates to the  
7 polycyclic aromatic hydrocarbon pyranthrene. The following four scientific  
8 publications are of interest:

9 "Solvent effect on  $^1\text{O}_2$  yield and the Mechanism of Polycyclic  
10 Hydrocarbon Triplet State Quenching by Oxygen", A.P. Darmanyan, Chemical  
11 Physics Letters, 1983, Vol. 96, 3, 383-9. This paper describes singlet oxygen  
12 production as a tool to study the effect of solvent on the interactions of triplet  
13 states in certain PAH's including Pyranthrene.

14 "Spectroscopic Investigation of Fluorescence Quenching Agents. Part  
15 IV.", J. C. Fetzer, Applied Spectroscopy, 1993, Vol. 47, 11, 1775-9. This  
16 paper describes the selective quenching of certain PAH's, including  
17 Pyranthrene, by Nitromethane as a means of classification.

18 "Spectroscopic Investigation of Fluorescence Quenching Agents. Part  
19 II", J. C. Fetzer, Applied Spectroscopy, 1992, Vol. 46, 8, 1260-5. This paper  
20 describes the selective quenching of certain PAH's, including Pyranthrene, by  
21 Nitromethane as a means of classification.

22 "The Triplet Energies of Azulene, b-Carotene, and Ferrocene", W.G.  
23 Herkstroeter, Journal of A.C.S., 1975, Vol. 97, 15, 4161-7. This paper  
24 describes the triplet energies of Azulene, b-Carotene, and Ferrocene by  
25 investigating their interaction with the triplet states of various PAHs including  
26 Pyranthrene.

27 Because of the problems that still prevail in the state-of-the-art of  
28 optical chemical sensors, need for improvement regarding several features of

1 these devices still exists. Specifically, a need exists to improve the stability of  
2 the indicator dye and matrices in order to improve their overall performance.  
3 A need also exists for a dye with a large extinction coefficient to obtain  
4 increased fluorescence intensity. A need also exists for longer wavelength  
5 dyes compatible with LED's. A need also exists for an improved dye with self  
6 referencing dual wavelength properties, both for excitation and emission.

7 In light of the foregoing, the principal objective of the present invention  
8 is to provide an improved optical chemical sensor system which overcomes  
9 the above-described shortcomings of the prior art and fulfills the above-listed  
10 needs.

## 11 SUMMARY OF THE INVENTION

12 The present invention provides a method and apparatus by which a  
13 single fluorescent dye is used both as the analyte indicator and as a reference  
14 element in an optical chemical sensor for sensing analytes such as oxygen,  
15 sulfur dioxide, nitric oxide, and many halogenated compounds which are  
16 known to quench the fluorescence of polycyclic aromatic hydrocarbons. In  
17 accordance with the invention the polycyclic aromatic hydrocarbon  
18 pyranthrene or the polycyclic heterocyclic aromatic compound flavanthrene,  
19 or a suitable derivative of these compounds is used as the single reference dye  
20 of the optical chemical sensor. This is possible because these dyes exhibit  
21 significantly different analyte quenching slopes at different excitation  
22 wavelengths, and also significantly different analyte quenching slopes at  
23 different emission wavelengths. The excitation wavelength of these dyes is  
24 advantageously in the range of the light generated by light emitting diodes  
25 and therefore the sensors utilizing these dyes can operate with LEDs.  
26 Advantageously these dyes also exhibit large Stokes shifts of emission  
27 wavelengths, and large extinction coefficients which result in strong  
28 fluorescence emissions. Excitation of one of these dyes at two or more



1 wavelengths, coordinated with the measurement of both the excitation  
2 intensities and one or more emission intensities, provides a self referencing  
3 system. Alternatively, excitation of these dyes at a single wavelength while  
4 measuring two or more emission intensities also provides a self referencing  
5 system. In combination with a suitable gas permeable matrix, such as  
6 polydimethylsiloxane, pyranthrene, flavanthrene or a suitable derivative of  
7 either of these two dyes provides improved means for the measurement and  
8 normalization of measurements of the analytes known to quench the  
9 fluorescence of polycyclic aromatic hydrocarbons.

10 In accordance with this invention, a hydrophobic oxygen (or other  
11 analyte) sensitive matrix is formed incorporating the pyranthrene, flavanthrene  
12 or suitable derivative of either of these two dyes. A suitable derivative is one  
13 that is sensitive to changes in oxygen (or other analyte) concentration in  
14 substantially the same manner as the parent compound, pyranthrene or  
15 flavanthrene. A preferred gas permeable matrix is polysiloxane polymer. The  
16 fluorescent pyranthrene, flavanthrene or suitable derivative dye dispersed or  
17 immobilized in the siloxane polymer matrix is coated at the end of a fiber, cast  
18 into films, or mounted with any of the other common methods for preparing  
19 optical sensors. For measurement the matrix with the dye is exposed to the  
20 analyte. The sensor apparatus in accordance with the invention further  
21 includes light sources, preferably and advantageously in accordance with the  
22 present invention one or two LEDs, preferably means for measuring the  
23 excitation light at one or two wavelengths, means for conducting the excitation  
24 light at one or two excitation wavelengths bands to the sensor dye, means for  
25 separating the fluorescent emission light at one or two wavelengths bands  
26 from the excitation light, means for measuring the emitted fluorescent light at  
27 one or two wavelengths bands, and means for calculating ratios of the  
28 measured intensities to arrive to the concentration of the analyte in a sample,

1 based on the spectral response curves of the pyranthrene, flavanthrene or  
2 derivative dye which is used in the sensor.

### 3 BRIEF DESCRIPTION OF THE DRAWINGS

4 **Figure 1** is a diagrammatic view of a first preferred embodiment of an  
5 optical chemical sensor apparatus in accordance with this invention;

6 **Figure 2** is a diagrammatic view of a second preferred embodiment of  
7 an optical chemical sensor apparatus in accordance with this invention;

8 **Figure 3** is an expanded view in section of an optical fiber optical  
9 chemical sensor in accordance with this invention;

10 **Figure 4** is a diagrammatic illustration of a test run using a sensor of  
11 this invention;

12 **Figures 5, 6, 7 and 8** are diagrammatic illustrations of the spectral  
13 response of a pyranthrene test cell;

14 **Figure 9** discloses the chemical structure of flavanthrene;

15 **Figure 10** discloses the chemical structure of pyranthrene;

16 **Figure 11** is a diagrammatic illustration of the spectral response of  
17 flavanthrene dye.

### 18 DETAILED DESCRIPTION OF THE INVENTION

19 The following specification taken in conjunction with the drawings sets  
20 forth the preferred embodiments of the present invention. The embodiments  
21 of the invention disclosed herein are the best modes contemplated by the  
22 inventors for carrying out their invention in a commercial environment,  
23 although it should be understood that various modifications can be  
24 accomplished within the parameters of the present invention.

25 The present invention is described below in connection with measuring  
26 oxygen concentration in an analyte fluid, because measuring oxygen, primarily  
27 in medical applications, is one of the most frequent use of optical chemical  
28 sensors, and the presently preferred embodiments of the invention are also

1 primarily directed to this use. It should be clearly understood however, that  
2 instead of oxygen other analytes of the type which quench the fluorescence of  
3 PAH dyes can also be measured in accordance with the present invention.

4 As is known in the prior art, in an optical chemical sensing apparatus  
5 for measuring partial pressure of oxygen ( $P_{O_2}$ ), a polycyclic aromatic  
6 hydrocarbon or polycyclic heterocyclic aromatic compound (PAH or HAC) is  
7 immobilized or dispersed in a gas permeable polymer such as polysiloxane.  
8 The PAH or HAC has an emission intensity which is inversely proportional to  
9 the oxygen concentration in accordance with the Stern-Volmer relationship.  
10 This relationship is described by equation (1):

$$11 \quad I_0/I = 1 + k_q t_0 [Q] \quad (1)$$

12 where  $I$  is the intensity,  $I_0$  is the intensity without quenching,  $k_q$  is the rate  
13 constant of the quenching reaction,  $t_0$  is the life of the emission in the absence  
14 of quencher, and  $[Q]$  is the concentration of quencher.  $t_0$  and  $k_q$  are constants  
15 which vary widely among PAH's.  $I_0$  is a constant which must be calibrated for  
16 a given optical system with sensor. In practical actual instruments in  
17 accordance with the state-of-the-art, equation (1) is typically further refined to  
18 accommodate certain offsets in the system, temperature coefficients for the  
19 various constants, and variations in the source light intensity as measured by a  
20 separate photodetector.

21 For the dyes of this invention where two or more wavelengths exhibit  
22 differing slopes, the equations become:

$$23 \quad I_{Oa}/I_a = 1 + k_{qa} t_{Oa} [Q] \quad (2)$$

$$24 \quad I_{Ob}/I_b = 1 + k_{qb} t_{Ob} [Q] \quad (3)$$

25 Where "a" and "b" in subscripts designate the values of " $I_0$ ", " $I$ ", " $q$ " and " $t$ "  
26 for the respective wavelength "a" and "b". In general,  $I_{Oa}/I_{Ob}$  is a constant for  
27 any given optical system independent of the sensor attached. These equations  
28 can then be solved for any random sensor attached to the optical system as

1 follows:

$$2 \quad I' = I_{Oa} / I_{Ob} \quad (4), \text{ and}$$

$$3 \quad [Q] = ((I_a / I_b) I') / ((k_{qa} t_{Oa}) + (k_{qb} t_{Ob}) (I_a / I_b)) \quad (5)$$

4  
5 Thus, in an optimum configuration the present invention, by utilizing  
6 pyranthrene or flavanthrene (or a suitable derivative of either of the two)  
7 enables the fabrication and use of a factory calibrated optical system which  
8 can be used interchangeably with any sensor without field calibration. The  
9 prior art has attempted to obtain this result through the use of two different  
10 dyes, but only with limited success.

11 As is described in more detail below, the pyranthrene or flavanthrene  
12 (or suitable derivative) dye can be, and preferably is covalently bonded to a  
13 siloxane polymer to provide a hydrophobic matrix from which the dye is not  
14 leachable.

15 The PAH and HAC dyes used in the prior art typically absorb light in  
16 the ultraviolet region of the spectra and emit at longer wavelengths, typically  
17 in the blue or green region of the spectrum. The pyranthrene or flavanthrene  
18 (or suitable derivative) dyes of this invention are among a very limited number  
19 of stable PAH's which excite in the blue region and have emission in the green  
20 region of the spectrum.

21 Whereas the fluorescence of certain PAHs and their dynamic quenching  
22 by oxygen and other analytes is well known, the present inventors are unaware  
23 of any prior description of fluorescence by flavanthrene, and are aware of only  
24 limited scientific literature (four articles) describing the fluorescent properties  
25 of pyranthrene. These are listed in the Brief Description of the Prior Art  
26 section of the present application for patent. None of the prior art pertaining  
27 to pyranthrene envisions this invention, which is based on the discovery of  
28 unusual quenching properties of pyranthrene and flavanthrene. Whereas other

1 PAHs and HACs tested show substantially uniform quenching across the  
2 entire excitation and emission spectrums, these indicators demonstrate  
3 substantially different quenching constants for different wavelengths in the  
4 excitation and emission spectra.

5 Thus, in one embodiment or preferred mode of the method and  
6 apparatus of the present invention two or more excitation wavelengths and a  
7 common emission wavelength are measured to provide a ratiometric intensity  
8 measurement. In this manner, by providing self-reference, the invention  
9 improves performance, increases stability, minimizes interference from  
10 environmental perturbations and optical system variations, and the like.

11 In another embodiment or preferred mode of the method and apparatus  
12 of the present invention a single excitation wavelength and two or more  
13 emission wavelengths are measured to provide a ratiometric intensity  
14 measurement, again providing a self-reference that improves performance,  
15 increases stability, minimizes interference from environmental perturbations  
16 and optical system variations, and the like.

17 In still another embodiment or preferred mode of the method and  
18 apparatus of the present invention a combination of multiple excitation and  
19 multiple emission wavelengths are measured to provide cross-referenced  
20 ratiometric intensity measurements, providing multiple self reference that  
21 improves performance, stability, and minimizes interferences. Of course it  
22 will be readily understood by those skilled in the art that reference to a  
23 "wavelength" in this description refers to a central wavelength and a band  
24 width corresponding thereto, which is often but not necessarily is  
25 characterized by its full width at half maximum intensity. The term  
26 "wavelength" thus refers to a band in this specification and claims, unless  
27 another meaning is apparent from context.

28 It has been discovered as still another unique feature of the present

1 invention that the pyranthrene and flavanthrene dyes have longer excitation  
2 wavelengths than most previously disclosed PAHs and HACs for use in  
3 chemical sensors. This makes it possible to use relatively inexpensive LEDs  
4 as sources of excitation light in the light sensing apparatus incorporating these  
5 dyes as sensors. A further uniqueness of this invention is the discovery that  
6 these dyes have a large extinction coefficient and strong fluorescence  
7 intensity.

8 A still further unique feature of the present invention is the discovery  
9 that the pyranthrene and flavanthrene dyes have a steeper oxygen response  
10 slope than most previously disclosed PAHs and HACs for use in similar  
11 optical chemical sensors. This makes the sensing devices incorporating these  
12 dyes more sensitive and therefore more accurate.

13 To form the actual sensor, the pyranthrene or flavanthrene dye or a  
14 suitable derivative of either one of these two dyes can be perfused in an  
15 insoluble matrix. However, it is preferred to immobilize the dye by covalently  
16 bonding it to the insoluble matrix, which is preferably a polysiloxane matrix.  
17 Alternatively, the pyranthrene or flavanthrene dye (or suitable derivative) can  
18 be perfused or immobilized in a fluoropolyurethane matrix, as is done with  
19 other agents in U. S. Patent No. 5,681,532, incorporated herein by reference.  
20 Generally speaking, chemical processes and methods normally used in the art  
21 for covalently bonding PAH compounds to a polysiloxane matrix can be  
22 utilized for bonding the pyranthrene or flavanthrene dye (or suitable  
23 derivative) to a polysiloxane matrix. For example, the methods of coupling  
24 described in U. S. Patent Nos. 4,746,751 and 5,015,715 can be used. U. S.  
25 Patent Nos. 4,746,751 and 5,015,715 are incorporated herein by reference.  
26 The methods and compositions utilized in the state-of-the-art for bonding PAH  
27 compounds to a polysiloxane matrix usually involve the use of linking  
28 moieties that covalantly link the PAH to the matrix through several atoms. A

1 "suitable derivative" of pyranthrene or flavanthrene within the scope of this  
2 invention is a chemically substituted pyranthrene or flavanthrene molecule  
3 which substantially retains one or more of the important fluorescent and  
4 quenching properties of the parent compound, pyranthrene or flavanthrene,  
5 respectively. In fact, a pyranthrene or flavanthrene moiety that is covalently  
6 linked to a matrix can also be viewed as a "pyranthrene or flavanthrene  
7 derivative", because any linking moiety that is attached to the pyranthrene or  
8 flavanthrene itself can be viewed as a substituent. As experience in the prior  
9 art with other PAH compounds used as optical sensors has demonstrated, the  
10 linking moieties that covalently bond PAHs to the matrix can be selected to  
11 avoid adverse effects on the fluorescent and quenching properties of the PAH  
12 compound.

13 A presently preferred method of using pyranthrene in an optical sensor  
14 is to link it to a polydimethylsiloxane methyl vinyl copolymer, which is  
15 thereafter cross-linked with polymethylhydrosiloxane as described below.

#### 16 Example 1: Preparation of Optical Chemical Sensor Matrix Having 17 Covalently Linked Pyranthrene

18 The chemical pyranthrene was obtained from (NBS Biologicals Ltd.,  
19 Huntingdon, U. K.). It can also be prepared in accordance with the chemical  
20 literature, see *E. Clar* Berichte Deutscher Chemischer Gesellschaft, 76, 332  
21 (1943).  $5 \times 10^{-4}$  mole pyranthrene and 5 grams of polydimethylsiloxane  
22 methyl vinyl copolymer PS925 (United Chemical Tech., Bristol Pa.) was  
23 measured into a 20 ml vial equipped with a stirrer. The solution was stirred  
24 and heated under nitrogen. Heating was continued until the temperature of the  
25 silicone reached 180 ° C, at which time 0.5 g of  $AlCl_3$  (Aldrich Chemical Co.)  
26 was added and the mixture was heated for an additional 15 minutes. The  
27 mixture was then cooled to room temperature and the  $AlCl_3$  was quenched  
28 with water. The organic phase was extracted with 50 ml of xylenes (Aldrich

1 Chemical Co.), washed twice with water, then dried over magnesium sulfate  
2 (Aldrich Chemical Co.). The xylenes were evaporated and the polysiloxane  
3 containing the covalently bonded pyranthrene was filtered through a glass  
4 wool plug. The dye modified PS925 polysiloxane was mixed 1:1 with  
5 polymethylhydrosiloxane PS119 (United Chemical Tech.) along with a trace  
6 of PC075 catalyst platinum, 1,3-diethenyl-1,1,3,3-tetramethyldisiloxane  
7 complexes in polydimethylsiloxane, vinyl terminated (United Chemical  
8 Tech.). One drop was applied to a test cell then cured at 55 ° C for 2 hours,  
9 resulting in a thin film ready for testing and use.

#### 10 Example 2: Preparation of Flavanthrene

11 Flavanthrene was prepared from flavanthrone (TCI America, Portland  
12 Oregon) by heating a mixture of flavanthrone (0.57 g) and 20.0 ml of 57% HI  
13 (Aldrich) to reflux for 20 hours. The hot solution was poured into 300 ml of  
14 2% aqueous sodium bisulfite solution (Sigma Chemical). The resulting  
15 precipitate was filtered, washed with cold water and air-dried. Further  
16 purification was carried out by adding the crude solid product to xylenes  
17 (Aldrich). Insoluble impurities were removed by filtration. The xylenes were  
18 removed by evaporation at 65 °C under reduced pressure. A dilute solution  
19 of approximately 0.01 mmol was prepared in toluene and placed in a 1 cm  
20 square test cell for testing.

21  
22 The excitation (EX) and emission (EM) spectra of the pyranthrene  
23 polysiloxane matrix and of the flavanthrene materials prepared above, were  
24 scanned with a Perkin-Elmer LS-5 spectrofluorimeter in the presence of  
25 oxygen (100% O<sub>2</sub>), room air (21% O<sub>2</sub>), and Nitrogen (0% O<sub>2</sub>).

26 The intensity measurements are set forth in **Table 1**.



TABLE 1

PAH	EX	EM	Oxygen	Room Air	Nitrogen	$I_{N_2}/I_{air}$	Scan #
Pyranthrene	410	441	28.9	60.6	122.9	2.03	a
Pyranthrene	410	500	27.2	47.5	79.3	1.67	b
Pyranthrene	456	500	74.8	144	204.3	1.42	c
Flavanthrene	410	444	62.4	120.1	198.4	1.65	d
Flavanthrene	410	472	106.9	209.3	299.2	1.43	e
		Scan #	Slope differential (%)				
		a vs. b	21.6				
		a vs. c	43				
		b vs. c	17.6				
		d vs. e	15.3				

Where Slope differential (a vs. b) =  $100 * ((I_{N_2}/I_{air})_a - (I_{N_2}/I_{air})_b) / (I_{N_2}/I_{air})_b$  and where  $I_{N_2}$  and  $I_{air}$  represent measured light intensities in the presence of nitrogen and air respectively, and the subscripts "a" and "b" refer to the measurements obtained in the corresponding scan designated "a" and "b".

Among all PAH compounds tested in connection with the present invention only pyranthrene and flavanthrene exhibited significant (>10%) dual wavelength variation in fluorescence (singlet) quenching coefficient. It should be readily apparent to those skilled in the art based on the foregoing data that the use of the single pyranthrene (or flavanthrene) indicator system is superior to the mixing of two different dyes. The superiority is due to the elimination of differential photodegradation, differential temperature coefficients,

1 differential chemical interferences, and the like.

2

### 3 Detailed Description of the Preferred Embodiments of the Optical 4 Sensing Apparatus of the Invention

5 Referring to the drawings which illustrate the preferred embodiments of  
6 the apparatus used in accordance with the invention, **Figure 1** illustrates a  
7 preferred form of the spectrofluorimetric system **10** for measurement of  
8 analyte in a continuous flow environment, or of samples of only a few  
9 microliters. The system includes a control section (CONTROLLER) for  
10 operation of multiple light sources (LAMP1) and (LAMP2), reading of the  
11 signals from photodetectors (PHOTODETECTOR 1, 2, 3), computations, and  
12 display handling. The CONTROLLER may be connected to a general purpose  
13 computer or may itself include a processing unit, and receives as input the data  
14 of measurements. The processing unit or computer calculates analyte  
15 concentration by using an algorithm that is based on the equations described  
16 above and includes as constants the previously observed characteristics of the  
17 indicator dye's response to the analyte at two excitation or two emission  
18 wavelengths or both.

19 The light sources may be in the form of LEDs, incandescent bulbs,  
20 lamps, or strobe lights. Light from each source (LAMP1, LAMP2) in  
21 sequence passes through a corresponding optical filter ( $\lambda_1$ ,  $\lambda_2$ ) for excitation  
22 wavelength and bandwidth selection, then enters two corresponding fibers  
23 (fibers 1 & 3, or fibers 2 & 4). Fibers 1 & 2 feed source light into Detector 1  
24 where the intensity of each lamp is measured in order to compensate for any  
25 variations. Fibers 3 and 4 feed source light into the multifiber bundle labeled  
26 fiber 7.

27 The distal end of fiber 7 joins with an optical analyte sensor fiber 8  
28 through an optical fiber connector (CONNECTOR). The source light reaches

1 the sensor tip (SENSOR) where it excites emission light from the indicator dye  
2 that is in accordance with the present invention, that is pyranthrene or  
3 flavanthrene or a suitable derivative, more preferably pyranthrene covalently  
4 linked to cross-linked polydimethylsiloxane methyl vinyl copolymer. The  
5 configuration of the sensor tip is shown in detail in Figure 3.

6 Fluorescent emission light returning from the optical analyte sensor  
7 material passes through fiber 8, back through the connector, and then into fiber  
8 7. The emission light is then split with a portion entering fibers 5 and 6. This  
9 emission light emerges from the proximal end of fibers 5 and 6, is focused  
10 with a lens, optically filtered for specific wavelengths by filters 13 and 14, then  
11 illuminates detectors 2 and 3 respectively. The electronic output of detectors  
12 1, 2 and 3 is connected to the controller for processing. In this embodiment  
13 the optical cable also acts as a beam splitter, although in other embodiments  
14 other devices may be used for this purpose. Photodetectors 2 and 3 may also  
15 be replaced with a suitable spectrometer for measurement of the entire  
16 emission spectrum.

17 Data processed by the control unit may be displayed digitally or used as  
18 the input to another computer system. It should be understood that the system  
19 of Figure 1 is for purposes of illustration only and that other processing and  
20 light control and transmission and receiving systems may be used.

21 **Figure 2** is a diagram of an alternative preferred embodiment used in  
22 preliminary optical sensor tests. The system includes a control section  
23 (CONTROLLER) for operation of multiple light sources (LED  $\lambda_1$ ) and (LED  
24  $\lambda_2$ ), reading of the signals from Detectors 1 & 2, computations, and display  
25 handling. LED  $\lambda_1$  (430nm) and LED  $\lambda_2$  (470nm) are alternatively illuminated  
26 to provide source light. This source light passes through a common excitation  
27 filter 14 where longer wavelengths are removed (470nm short pass) and into a  
28 glass optical window. Attached to the optical window with an index matching

1 optical fluid or adhesive is a clear plastic substrate with indicator dye sensor  
2 film, containing the pyranthrene or flavanthrene dye in accordance with the  
3 invention, preferably pyranthrene covalently linked to cross-linked  
4 polydimethylsiloxane methyl vinyl copolymer. The source light passes from  
5 the optical window through the substrate and into the sensor film.

6 A portion of the emission light resulting from indicator dye excitation  
7 by the source light passes back through the substrate, through the optical  
8 window, and then through emission cut-off filter  $\lambda 3$ . The emission filter  
9 (480nm long pass) removes any scattered excitation light from the returning  
10 light and allows the remainder to pass into Detector 2 for measurement. This  
11 embodiment represents a disposable sensor element without the requirement  
12 for a costly fiber optic connector. Simple coated films may be used.

13 **Figure 3** illustrates the end of an optical chemical sensor structure 120  
14 in accordance with this invention. In this embodiment the front face 122 of  
15 the optical fiber 128 is coated with the indicator matrix 125 using the  
16 immobilized pyranthrene materials previously described, the analyte sensor  
17 then being overcoated with an opaque analyte permeable layer 127 (Black  
18 RTV silicone GE 283).

19 The physical structure of the optical chemical sensors in accordance  
20 with this invention may vary widely because of the ease of use, stability, and  
21 relatively high fluorescence intensity. The diameter of the optical fiber in a  
22 sensor may vary from a few microns to 1/2 inch or more. One of the  
23 advantages of the present invention is the ability to provide an optical  
24 chemical sensor in which the diameter is quite small and easily inserted  
25 invasively in a hypodermic syringe or other form of invasive catheter. In the  
26 case of coating(s) on the end face of the optical fiber, films in the range of 20  
27 to 40 microns are preferred, although films greater than 300 microns have also  
28 been found to be effective with somewhat slower response.

1           The optical fiber may be of plastic material or of glass. The term  
2 optical fiber as used in this invention also includes optical windows and rods  
3 which may be used in sensors without cladding.

4           **Figure 4** charts the oxygen response of the signal ratio for a sensor  
5 fabricated in accordance with this invention using pyranthrene immobilized in  
6 silicone.

7           **Figure 5** illustrates the fluorescent emission response of a pyranthrene  
8 polysiloxane polymer at 500 nm when excited in a range of excitation  
9 wavelengths between 260 and 490 nm in the presence of varying  
10 concentrations of oxygen. Thus, by way of example and not a limitation on  
11 this invention, curve 65 represents the excitation spectrum of the sensor  
12 material at 0% oxygen. Curve 65a represents the response at 20.3% oxygen,  
13 and curve 65b represents the response at 100% oxygen.

14           **Figure 6** charts the emission ratios at 500 nm when the excitation is at  
15 412nm and at 457nm in the presence of varying concentrations of oxygen.  
16 This illustrates one computational method for self compensation in the  
17 measurement of oxygen.

18           **Figure 7** illustrates the emission spectral response of a pyranthrene  
19 sensor when excited at 410nm. Thus, by way of example and not a limitation  
20 on this invention, curve 67 represents the excitation spectrum of the sensor  
21 material at 0% oxygen. Curve 67a represents the response at 20.3% oxygen,  
22 and curve 67b represents the response at 100% oxygen.

23           **Figure 8** charts the ratio of the 442nm and 470nm emissions, when  
24 excited at 412 nm in the presence of varying concentrations of oxygen. This  
25 illustrates another computational method for self compensation in the  
26 measurement of oxygen.

27           **Figure 9** discloses the chemical structure of the flavanthrene molecule.

28           **Figure 10** discloses the chemical structure of the pyranthrene molecule.

1           **Figure 11** illustrates the emission spectral response of flavanthrene in  
2 toluene when excited at 410nm in the presence of varying concentrations of  
3 oxygen. Thus, by way of example and not a limitation on this invention, curve  
4 69 represents the excitation spectrum of the sensor material at 0% oxygen.  
5 Curve 69a represents the response at 20.3% oxygen, and curve 69b represents  
6 the response at 100% oxygen.

7           Optical chemical sensor systems or apparatus can be constructed in  
8 accordance with the present invention to measure concentrations of oxygen,  
9 nitric oxide, sulfur dioxide, or halogens, or other analytes known to quench  
10 the fluorescence of polycyclic aromatic hydrocarbons. Sensors constructed in  
11 accordance with the present invention can also be used for measuring glucose  
12 by means of an enzymatically coupled system, and aerodynamic pressure by  
13 means of an oxygen concentration coupled system. Moreover, optical  
14 chemical sensor systems or apparatus can also be constructed in accordance  
15 with the present invention where the initial excitation is provided by applying  
16 an electric field to the probe containing the dispersed or immobilized  
17 pyranthrane, flavanthrane (or suitable derivative) dye.

18           What has been described above is an improved self-referencing optical  
19 chemical sensor system which overcomes several deficiencies of the prior art  
20 systems. It should be understood that while several embodiments of the  
21 invention have been shown and described, modifications may be become  
22 readily apparent to those skilled in the art in light of the foregoing disclosure.  
23 Therefore, the scope of the present invention should be interpreted solely from  
24 the following claims, as such claims are read in light of the disclosure.

## WHAT IS CLAIMED IS:

- 1  
2           1.     An optical chemical sensor system for determining the  
3 concentration of an analyte, comprising:  
4           a probe having a polymeric matrix and an optical sensing compound  
5 dispersed or immobilized in the matrix, the sensing compound being  
6 responsive to excitation by light and providing a fluorescent emission of light  
7 in response to the excitation, said fluorescence emission being quenched by  
8 the analyte in a relationship that is a function of the concentration of the  
9 analyte, the sensing compound being selected from the group consisting of  
10 pyranthrene, a derivative of pyranthrene, flavanthrene and a derivative of  
11 flavanthrene;  
12           a source of light that provides excitation to the optical sensing  
13 compound in the probe;  
14           a detector that measures the intensity of the fluorescent emission of the  
15 sensing compound in the probe, and  
16           means utilizing an algorithm for calculating the concentration of the  
17 analyte from the measured intensity of the fluorescent emission, said algorithm  
18 being based on the spectral response curve of the optical sensing compound.  
19           2.     The optical chemical sensor system in accordance with Claim 1,  
20 where the source of light provides excitation at at least one predetermined  
21 wavelength and where the detector measures the intensity of the fluorescent  
22 emission of the sensing compound in the probe at at least two wavelengths of  
23 fluorescence emission.  
24           3.     The optical chemical sensor system in accordance with Claim 1,  
25 where the source of light provides excitation at at least two predetermined  
26 wavelengths and where the detector measures the intensity of the fluorescent  
27 emission of the sensing compound in the probe in response to each of the two  
28 excitation wavelengths independently.

1           4.     The optical chemical sensor system in accordance with Claim 2,  
2     where the sensing compound is pyranthrene.

3           5.     The optical chemical sensor system in accordance with Claim 4  
4     where pyranthrene is immobilized in the matrix by being covalently linked to  
5     the matrix.

6           6.     The optical chemical sensor system in accordance with Claim 3,  
7     where the sensing compound is pyranthrene.

8           7.     The optical chemical sensor system in accordance with Claim 6  
9     where pyranthrene is immobilized in the matrix by being covalently linked to  
10    the matrix.

11          8.     An optical chemical sensor system for determining the  
12    concentration of an analyte, comprising:  
13          a probe having a polymeric matrix and an optical sensing compound  
14    dispersed or immobilized in the matrix, the optical sensing compound being  
15    selected from the group consisting of pyranthrene, a derivative of pyranthrene,  
16    flavanthrene and a derivative of flavanthrene, and having the property of  
17    emitting fluorescence in response to excitation by light, said fluorescence  
18    being quenched by the presence of an analyte resulting in a change of  
19    fluorescence intensity depending on the concentration of the analyte, said  
20    fluorescent intensity being additionally dependent on the excitation  
21    wavelength and on the fluorescent emission wavelength, and being affected  
22    differently by the analyte at different excitation or emission wavelengths;  
23          light source means for providing excitation light at at least one  
24    wavelength;

25          means optically coupled to the light source means for defining a first  
26    optical path for the excitation light from the light source means to the optical  
27    sensing compound in the matrix of the probe;

28          means optically coupled to the optical sensing compound in the matrix



1 of the probe for defining a second optical path for fluorescent emission light  
2 from the optical sensing compound in the matrix of the probe;

3 detection means optically coupled to said second optical path for  
4 measuring emission light intensity at at least two different wavelengths from  
5 the optical sensing compound in the matrix of the probe;

6 computing means coupled to the detection means and receiving input  
7 from the detection means for calculating the concentration of the analyte to  
8 which the probe is exposed from the measurements of the emission light at at  
9 least two wavelengths of fluorescent emission, said calculation utilizing an  
10 algorithm based on the spectral response curve of the optical sensing  
11 compound.

12 9. The optical chemical sensor system in accordance with Claim 8  
13 wherein the optical sensing compound is pyranthrene.

14 10. The optical chemical sensor system in accordance with Claim 9  
15 wherein the matrix is a polysiloxane matrix to which the pyranthrene molecule  
16 is covalently linked.

17 11. The optical chemical sensor system in accordance with Claim 8  
18 adapted for determining the concentration of oxygen.

19 12. An optical chemical sensor system for determining the  
20 concentration of an analyte, comprising:

21 a probe having a polymeric matrix and an optical sensing compound  
22 dispersed or immobilized in the matrix, the optical sensing compound being  
23 selected from the group consisting of pyranthrene, a derivative of pyranthrene,  
24 flavanthrene and a derivative of flavanthrene, and having the property of  
25 emitting fluorescence in response to excitation by light, said fluorescence  
26 being quenched by the presence of an analyte resulting in a change of  
27 fluorescence intensity depending on the concentration of the analyte, said  
28 fluorescent intensity being additionally dependent on the excitation

1 wavelength and on fluorescent emission wavelength, and being affected  
2 differently by the analyte at different excitation and emission wavelengths;

3 light source means for providing excitation light at at least two  
4 wavelengths;

5 means optically coupled to the light source means for defining a first  
6 optical path for the excitation light from the light source means to the optical  
7 sensing compound in the matrix of the probe;

8 means optically coupled to the optical sensing compound in the matrix  
9 of the probe for defining a second optical path for fluorescent emission light  
10 from the optical sensing compound in the matrix of the probe;

11 detection means optically coupled to said second optical path for  
12 measuring fluorescent emission light from the optical sensing compound in the  
13 matrix of the probe at at least one wavelength in response to excitation at at  
14 least two wavelengths;

15 computing means coupled to the detection means and receiving input  
16 from the detection means for calculating the concentration of the analyte to  
17 which the probe is exposed from the measurements of the fluorescent emission  
18 light responsive to the excitation at each of the two wavelengths, said  
19 calculation utilizing an algorithm based on the spectral response curve of the  
20 optical sensing compound.

21 13. The optical chemical sensor system in accordance with Claim 12  
22 wherein the optical sensing compound is pyranthrene.

23 14. The optical chemical sensor system in accordance with Claim 13  
24 wherein the matrix is a polysiloxane matrix to which the pyranthrene molecule  
25 is covalently linked.

26 15. The optical chemical sensor system in accordance with Claim 8  
27 adapted for determining the concentration of oxygen.

28 16. An improvement in an optical chemical sensor system for

1 determining the concentration of an analyte, the improvement comprising:  
2 a probe containing an analyte sensor material which comprises a gas  
3 permeable composition containing fluorescent analyte sensitive indicator  
4 molecules selected from the group consisting of pyranthrene, a derivative of  
5 pyranthrene, flavanthrene and a derivative of flavanthrene.

6 17. The improvement defined in Claim 16, wherein:  
7 said analyte sensor material responds with a change in fluorescence intensity  
8 depending on the concentration of the analyte, and wherein the system  
9 comprises:

10 a source of excitation light;  
11 means optically coupled to said source for defining a first optical path  
12 for the excitation light from said source to said analyte sensor material;  
13 means optically coupled to said analyte sensor material for defining a  
14 second optical path for emission light from said analyte sensor material;  
15 a detection system optically coupled to said second optical path for  
16 detecting emission light from said analyte sensor material, and  
17 means coupled to said detection system for providing an output reading  
18 from said detection system.

19 18. The improvement defined in Claim 17 wherein:  
20 said chemical sensor system excites the analyte sensor material at a first  
21 excitation wavelength (EX1) which is responsive with a change in  
22 fluorescence intensity depending on the concentration of the analyte, and then  
23 at one or more additional excitation wavelengths (EX2) which are responsive  
24 with a significantly different change in fluorescence intensity depending on the  
25 concentration of the analyte, in which said detection system detects emission  
26 light at a wavelength (EM1) from said material for (EX1) and (EX2)  
27 independently, and in which said means for providing an output reading from  
28 said detection system measures the relative detected emission intensities at

1 (EM1) from said analyte sensor material.

2           **19.** The improvement defined in Claim 17 wherein:

3 said chemical sensor system further excites the analyte sensor material at a  
4 first excitation wavelength (E1), and in which said detection system detects  
5 emission light from said material at a first emission wavelength (EM1)  
6 independently which is responsive with a change in fluorescence intensity  
7 depending on the concentration of the analyte, and at one or more additional  
8 emission wavelengths (EM 2) which are responsive with significantly different  
9 changes in fluorescence intensity depending on the concentration of the  
10 analyte, and in which said means for providing an output reading from said  
11 detection system measures the relative detected emission intensities EM1 and  
12 EM2 from said chemical sensor material.

13           **20.** The improvement defined in Claim 16 where the analyte  
14 concentration to be determined is selected from the group consisting of  
15 oxygen, nitric oxide, sulfur dioxide, and halogens known to quench the  
16 fluorescence of polycyclic aromatic hydrocarbons.

17           **21.** The improvement defined in Claim 16 where the analyte  
18 concentration determined is glucose by means of an enzymatically coupled  
19 system.

20           **22.** The improvement defined in Claim 16 where the analyte  
21 concentration determined is aerodynamic pressure by means of an oxygen  
22 concentration coupled system.

23           **23.** The improvement defined in Claim 16, wherein:  
24 said analyte sensor material responds with a change in fluorescence intensity  
25 depending on the concentration of the analyte, and which includes: a source of  
26 excitation electric field, means electrically coupled to said source for applying  
27 the electric field across said analyte sensor material, means optically coupled  
28 to said analyte sensor material for defining an optical path for emission light

1 from said analyte sensor material, a detection system optically coupled to said  
2 optical path for detecting emission light from said analyte sensor material, and  
3 means coupled to said detection system for providing an output reading from  
4 said detection system.

5       **24.** A method of measuring the concentration of an analyte with an  
6 optical chemical sensing probe including an optical sensing compound  
7 dispersed or immobilized in a matrix included in the probe, the sensing  
8 compound being selected from the group consisting of pyranthrene, a  
9 derivative of pyranthrene, flavanthrene and a derivative of flavanthrene, the  
10 method comprising the steps of:

11       contacting the sensing compound included in the matrix with a fluid in  
12 which the analyte is to be measured;

13       providing an excitation light to the sensing compound in the matrix of  
14 the probe and in contact with the fluid;

15       measuring the intensity of a fluorescent emission light generated by the  
16 sensing compound in the matrix of the probe and in contact with the fluid;

17       computing the concentration of the analyte in the fluid from the  
18 measured fluorescent emission light, said step of computing utilizing an  
19 algorithm based on the spectral response curve of the optical sensing  
20 compound in the presence of varying concentrations of the analyte.

21       **25.** The method in accordance with Claim 24 where

22       the step of providing an excitation light to the sensing compound in the  
23 matrix of the probe and in contact with the fluid comprises providing  
24 excitation light at least two different wavelengths;

25       the step of measuring the intensity of a fluorescent emission light  
26 generated by the sensing compound in the matrix of the probe and in contact  
27 with the fluid comprises measuring the intensity of said fluorescent emission  
28 at a predetermined wavelength in response to each of said two excitations at

1 the two different excitation wavelengths, and

2 where said algorithm used in said step of computing is based on the  
3 spectral response curve at the two different wavelengths of excitation of the  
4 optical sensing compound in the presence of varying concentrations of the  
5 analyte.

6 26. The method in accordance with Claim 25 where the wavelengths  
7 of excitation is above 400 nanometers.

8 27. The method in accordance with Claim 26 where the analyte is  
9 selected from the group consisting of oxygen, nitric oxide, sulfur dioxide, and  
10 halogens known to quench the fluorescence of polycyclic aromatic  
11 hydrocarbons and glucose.

12 28. The method in accordance with Claim 26 where the analyte is  
13 oxygen.

14 29. The method in accordance with Claim 26 where the sensing  
15 compound is pyranthrane covalently linked to the matrix.

16 30. The method in accordance with Claim 24 where  
17 the step of measuring the intensity of a fluorescent emission light  
18 generated by the sensing compound in the matrix of the probe and in contact  
19 with the fluid comprises measuring the intensity of said fluorescent emission  
20 at at least two different and predetermined wavelengths in response to an  
21 excitation at a single wavelength, and

22 where said algorithm used in said step of computing is based on the  
23 spectral curve of the optical sensing compound in the presence of varying  
24 concentrations of the analyte at the two different wavelengths of fluorescent  
25 emission in response to the excitation at the single wavelength.

26 31. The method in accordance with Claim 30 where the wavelength  
27 of excitation is above 400 nanometers.

28 32. The method in accordance with Claim 31 where the analyte is

1 selected from the group consisting of oxygen, nitric oxide, sulfur dioxide, and  
2 halogens known to quench the fluorescence of polycyclic aromatic  
3 hydrocarbons and glucose.

4       33. The method in accordance with Claim 32 where the analyte is  
5 oxygen.

6       34. The method in accordance with Claim 31 where the sensing  
7 compound is pyranthrane covalently linked to the matrix.

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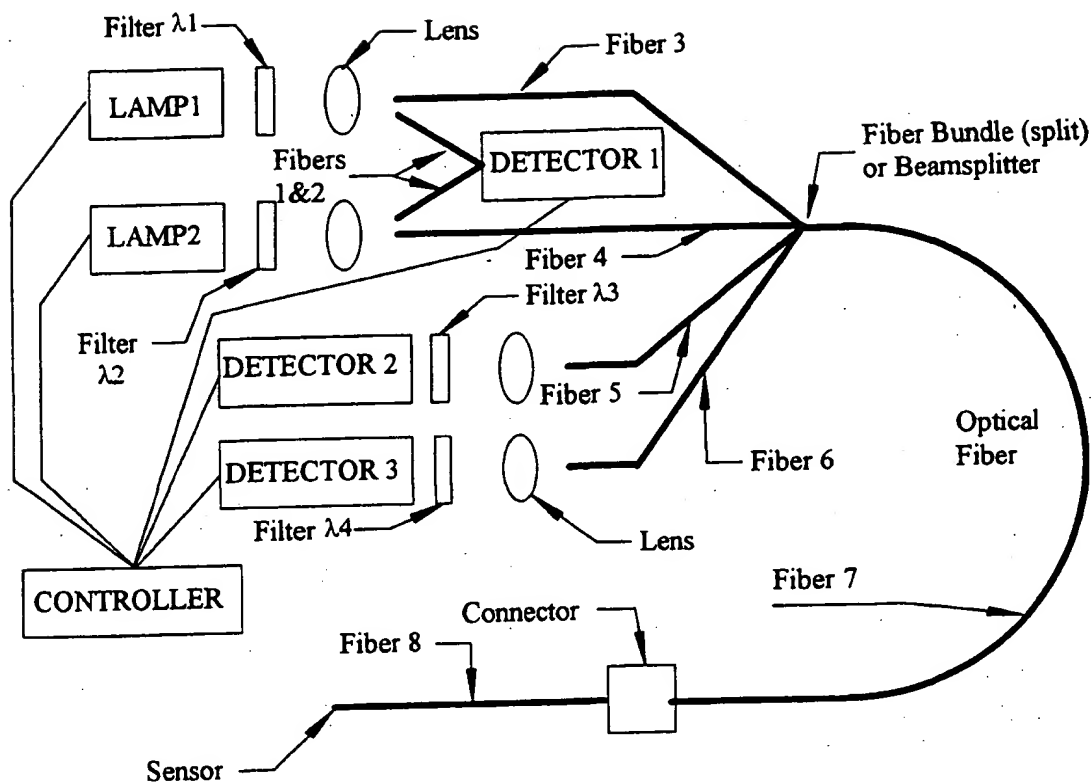


Figure 1

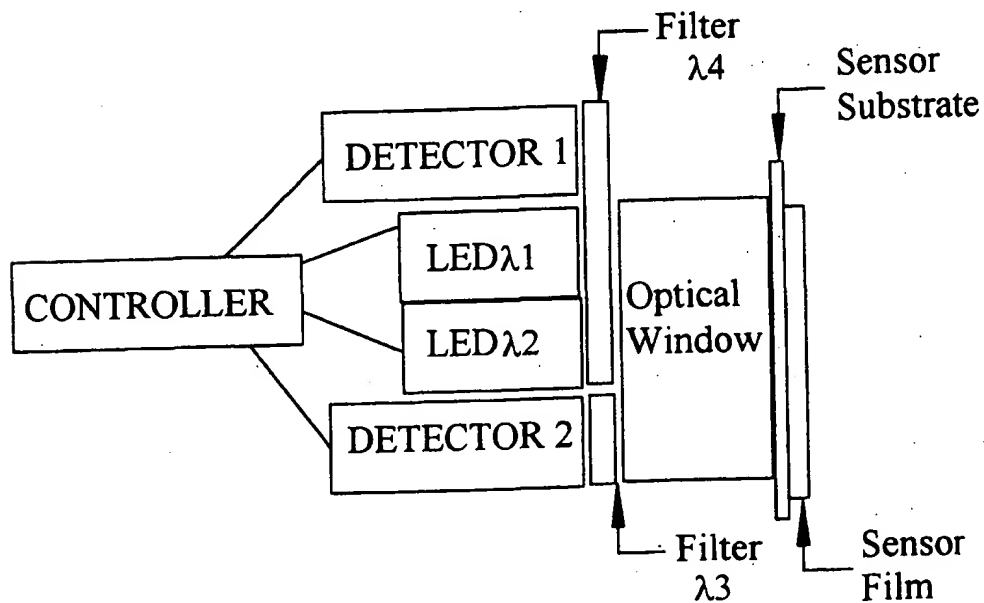


Figure 2



2/6

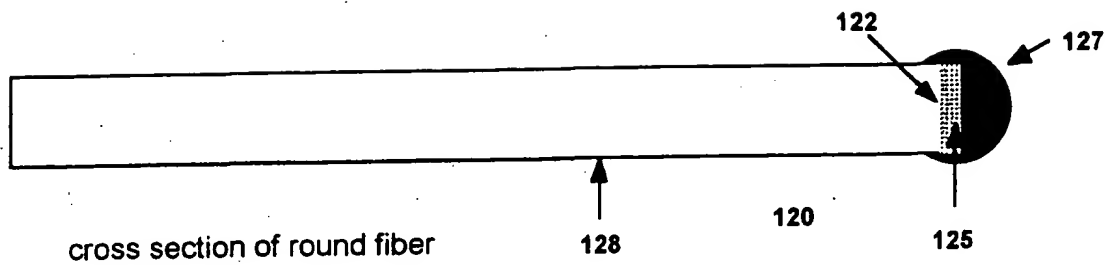


Figure 3

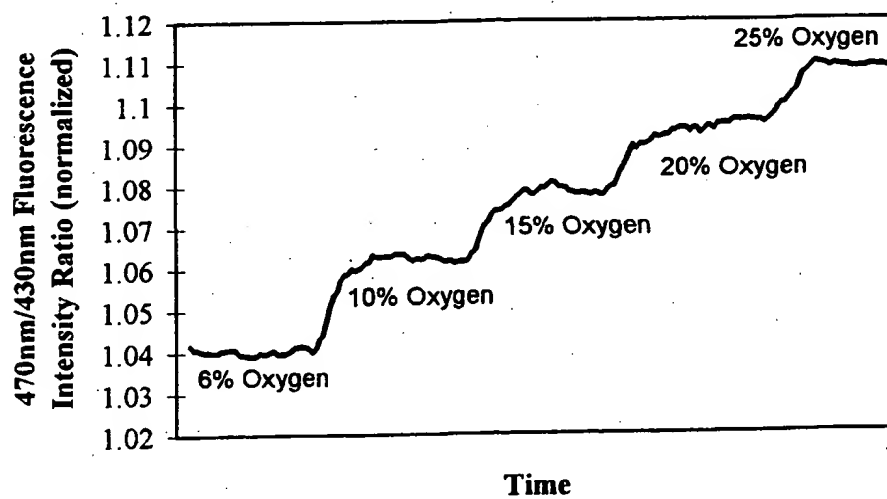
**Pyranthrene test run**

Figure 4

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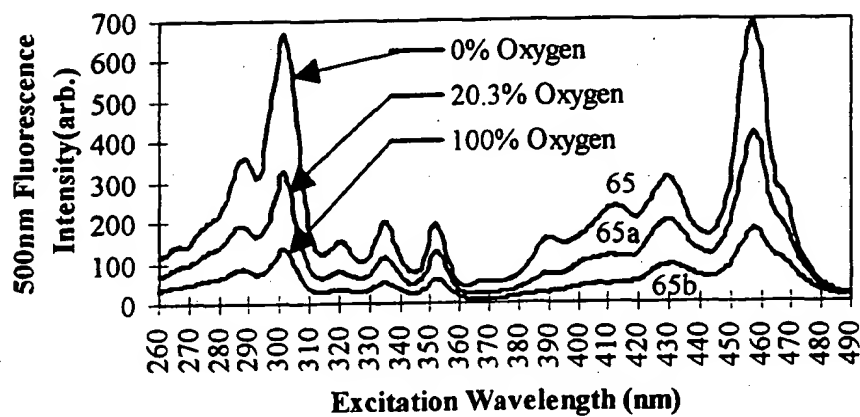
**Excitation Spectrum**

Figure 5

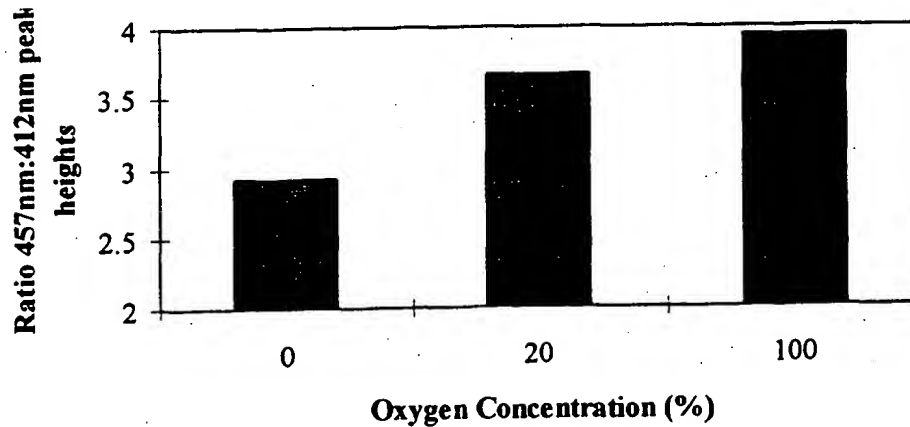
**Excitation Peak Ratios**

Figure 6

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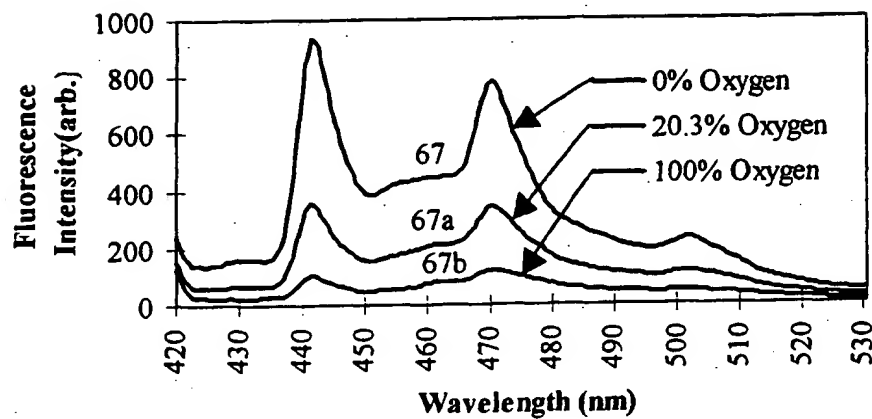
**Emission Spectra (excited at 412nm)**

Figure 7

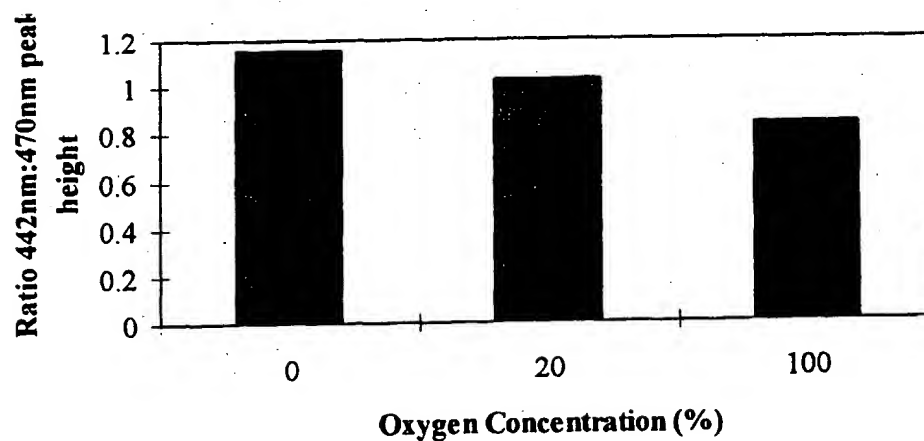
**Ratio of Emission Peak Heights**

Figure 8

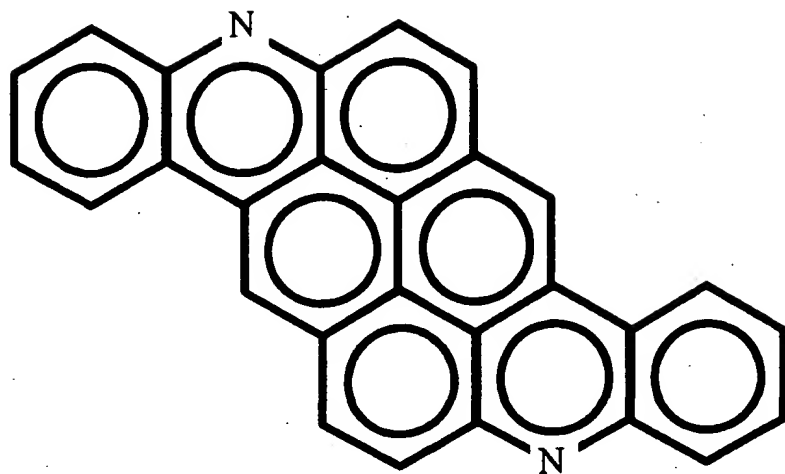


Figure 9

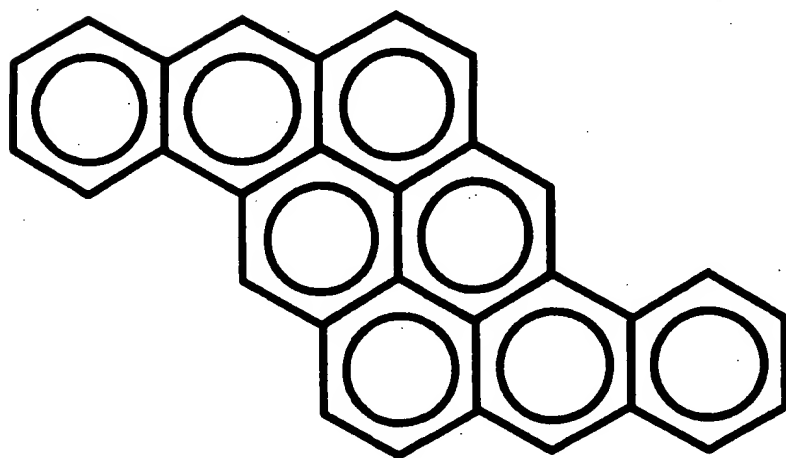


Figure 10

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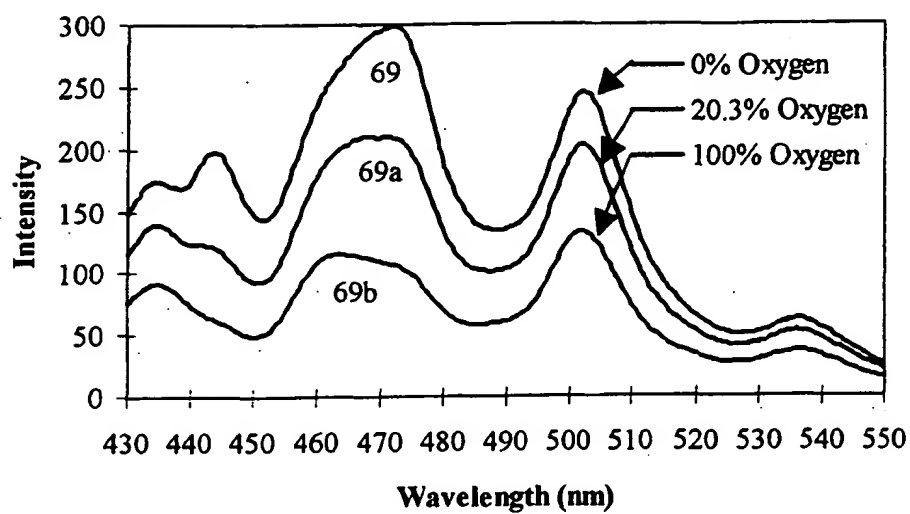


Figure 11

# INTERNATIONAL SEARCH REPORT

International Application No

PCT/US 00/00800

## A. CLASSIFICATION OF SUBJECT MATTER

IPC 7 G01N21/64 G01N21/77

According to International Patent Classification (IPC) or to both national classification and IPC

## B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 7 G01N

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

## C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	US 5 094 959 A (GOTTLIEB AMOS ET AL) 10 March 1992 (1992-03-10) cited in the application	1,2,5,7, 8,10,11, 14,15, 17,19, 20,24, 27,28, 30,32,33
A	column 3, line 42 -column 4, line 51 claims 1,2; figure 1 --- WO 81 01883 A (ELECTRO NUCLEONICS) 9 July 1981 (1981-07-09)  page 30 page 34 claim 1; table 4 --- -/-	1,4,6,8, 9,12,13, 16,24, 29,34



Further documents are listed in the continuation of box C.



Patent family members are listed in annex.

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Date of the actual completion of the international search

14 June 2000

Date of mailing of the international search report

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# INTERNATIONAL SEARCH REPORT

International Application No

PCT/US 00/00800

## C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	<p>US 5 681 532 A (KANE JAMES ET AL)  28 October 1997 (1997-10-28)  cited in the application</p> <p>column 11, line 5 - line 35  claim 1; example 4  -----</p>	<p>1,8,10,  11,  14-16,  20,24,  26-28,  31-33</p>

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Information on patent family members

International Application No

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